Solution

CHEMISTRY

MHT - CET - Chemistry

1.

(c) mole
Explanation:
mole

2.

(b) 6
Explanation:
6

3.

(d) square planar
Explanation:
square planar

4. **(a)** +5
Explanation:
+5

5.

(b) $2Fe^{2+} + 2H^{+} + H_2O_2 \rightarrow 2Fe^{3+} + 2H_2O$
Explanation:
$2Fe^{2+} + 2H^{+} + H_2O_2 \rightarrow 2Fe^{3+} + 2H_2O$

6.

(d) dipole-dipole
Explanation:
dipole-dipole

7.

(c) Option (C)
Explanation:
Option (C)

8. **(a)** Acetylene
Explanation:
Acetylene

9.

(c) $\dot{C}H_3$
Explanation:
$\dot{C}H_3$

10.

(c) 5×10^{23}
Explanation:

5×10^{23}

11. **(a)** Carbon dioxide
Explanation:
Carbon dioxide

12. **(d)** 8
Explanation:
8

13. **(b)** ΔT_b
Explanation:
ΔT_b

14. **(c)** Carbonated water
Explanation:
Carbonated water

15. **(a)** constant temperature
Explanation:
constant temperature

16. **(c)** energy changes in a system
Explanation:
energy changes in a system

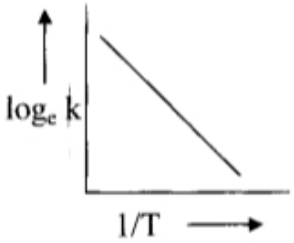
17. **(a)** $CaCO_{3(s)} \rightarrow CaO_{(s)} + CO_{2(g)}$
Explanation:
$CaCO_{3(s)} \rightarrow CaO_{(s)} + CO_{2(g)}$

18. **(b)** option (b)
Explanation:
ΔS is negative and therefore, ΔH should be highly negative.

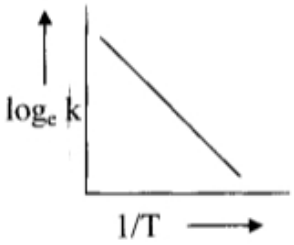
19. **(d)** $Q_P = \Delta U + P_{ext}\,\Delta V$
Explanation:
$Q_P = \Delta U + P_{ext}\,\Delta V$

20. **(c)** the minimum kinetic energy that the colliding reactant molecules must possess
Explanation:
the minimum kinetic energy that the colliding reactant molecules must possess

21.

(d) 

Explanation:



22.

(d) HF

Explanation:

HF

23.

(c) HOCIO

Explanation:

HOCIO

24.

(b) Square pyramidal

Explanation:

Square pyramidal

25.

(b) 4 - 7

Explanation:

4 - 7

26.

(b) $FeCr_2O_4$

Explanation:

$FeCr_2O_4$

27.

(b) II, IV

Explanation:

II, IV

28.

(d) Fe^{3+}, Mn^{2+}

Explanation:

Fe^{3+}, Mn^{2+}

29.

(b) $[Cr(H_2O)_4Cl_2]Cl.2H_2O$

Explanation:

$[Cr(H_2O)_4Cl_2]Cl.2H_2O$

30. **(a)** 5

Explanation:

5

31. **(a)** Dichlorobis(ethylenediamine)cobalt (III) ion

Explanation:

Dichlorobis(ethylenediamine)cobalt (III) ion

32.

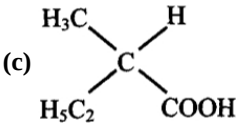
(c) isothiocyanato

Explanation:

isothiocyanato

33.

(c)



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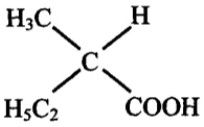


34.

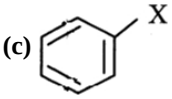
(d) 1-Bromo-1-phenylethane

Explanation:

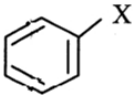
1-Bromo-1-phenylethane

35.

(c)



Explanation:



36. **(a)** Option (b)

Explanation:

- I effect of Cl disperses - ve charge on O atom to produce more stable anion

37.

(d) ethenol

Explanation:

ethenol

38. **(a)** $LiAlH_4$

Explanation:

$LiAlH_4$

39.

(b) methyl phenyl ether

Explanation:

methyl phenyl ether

40.

(c) Formaldehyde

Explanation:

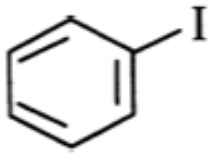
Formaldehyde

41. **(a)** H - COOH

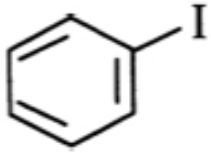
Explanation:

H - COOH

42.

(d) 

Explanation:



43.

(b) i - d, ii - c, iii - b, iv - e

Explanation:

i - d, ii - c, iii - b, iv - e

44.

(c) N,N-dimethyl-2-methylpropan-2-amine

Explanation:

N,N-dimethyl-2-methylpropan-2-amine

45.

(c) -NC

Explanation:

-NC

46.

(d) aldehydic and alcoholic

Explanation:

aldehydic and alcoholic

47.

(c) -CHO group

Explanation:

-CHO group

48. **(a)** Dacron

Explanation:

Dacron

49.

(b) condensation polymer

Explanation:

condensation polymer

50. **(a)** I, II and III

Explanation:

I, II and III